

Name:

Single Replacement Reactions Worksheet

Predict the products and balance the following single replacement reactions. If no reaction occurs write N.R.

For transition metals use the following charges:

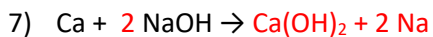
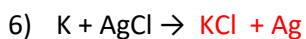
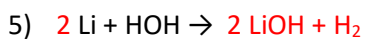
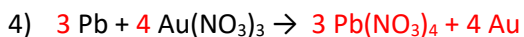
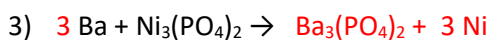
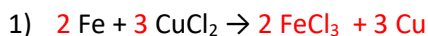
Iron: Fe³⁺

Mercury: Hg²⁺

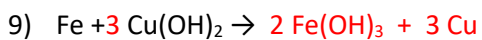
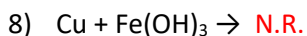
Lead: Pb⁴⁺

Copper: Cu¹⁺

Gold: Au³⁺



(Note- if this were to occur in water, which it should, the water is a reactant and reaction is a lie)



Write the reactions and predict the products of each of the following single replacement reactions. Balance all reactions. If no reaction occurs write N.R.

For transition metals use the following charges:

Iron: Fe^{3+}

Mercury: Hg^{2+}

Lead: Pb^{4+}

Copper: Cu^{1+}

Gold: Au^{3+}

- 1) Lead II chloride + magnesium



- 2) Barium nitrate + zinc



- 3) Potassium + tin IV nitrate



(Note- if this were to occur in water, which it should, the water is a reactant and reaction is a lie)

- 4) Copper + silver nitrate

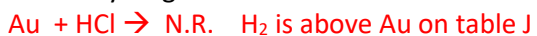


- 5) Sodium phosphate + potassium



(Note- if this were to occur in water, which it should, the water is a reactant and reaction is a lie)

- 6) Gold + hydrogen chloride



- 7) Magnesium + aluminum hydroxide



- 8) Iron + copper II sulfate



- 9) Iron + nickel II iodide



- 10) Sodium permanganate + calcium



(Note- if this were to occur in water, which it should, the water is a reactant and reaction is a lie)

- 11) Hydrogen chloride + zinc



- 12) Aluminum + iron II dichromate

