

- What is the oxidation number of hydrogen in  $\text{CaH}_2$ ?  
A) +1 B) +2 C) -1 D) -2
- The oxidation number of nitrogen in  $\text{N}_2$  is  
A) +1 B) 0 C) +3 D) -3
- In which compound does hydrogen have an oxidation state of -1?  
A)  $\text{NaH}$  B)  $\text{HCl}$  C)  $\text{H}_2\text{O}$  D)  $\text{NH}_3$
- What is the oxidation number of nitrogen in  $\text{HNO}_2$ ?  
A) -1 B) -2 C) +3 D) +4
- Oxygen has an oxidation number of -2 in  
A)  $\text{O}_2$  B)  $\text{NO}_2$   
C)  $\text{Na}_2\text{O}_2$  D)  $\text{OF}_2$
- What is the oxidation state of nitrogen in the compound  $\text{NH}_4\text{Br}$ ?  
A) -1 B) +2 C) -3 D) +4
- What is the oxidation number assigned to manganese in  $\text{KMnO}_4$ ?  
A) +7 B) +2 C) +3 D) +4
- What is the oxidation state of nitrogen in  $\text{NaNO}_2$ ?  
A) +1 B) +2 C) +3 D) +4
- What is the oxidation number of carbon in  $\text{NaHCO}_3$ ?  
A) -2 B) +2 C) -4 D) +4
- Chlorine has an oxidation state of +3 in the compound  
A)  $\text{HClO}$  B)  $\text{HClO}_2$   
C)  $\text{HClO}_3$  D)  $\text{HClO}_4$
- What is the oxidation number of sulfur in  $\text{H}_2\text{SO}_4$ ?  
A) 0 B) -2 C) +6 D) +4
- What is the oxidation state of phosphorus in the compound  $\text{Na}_3\text{PO}_3$ ?  
A) 0 B) -3 C) +3 D) +5
- Oxygen will have a positive oxidation number when combined with  
A) fluorine B) chlorine  
C) bromine D) iodine
- In which substance does sulfur have a negative oxidation number?  
A)  $\text{Na}_2\text{S}$  B)  $\text{CaSO}_4$   
C)  $\text{S}$  D)  $\text{SO}_2$
- In which compound does hydrogen have a negative oxidation number?  
A)  $\text{CaH}_2$  B)  $\text{H}_3\text{PO}_4$   
C)  $\text{NaOH}$  D)  $\text{NH}_3$
- If element  $X$  forms the oxides  $\text{XO}$  and  $\text{X}_2\text{O}_3$ , the oxidation numbers of element  $X$  are  
A) +1 and +2 B) +2 and +3  
C) +1 and +3 D) +2 and +4
- What is the oxidation state for a Mn atom?  
A) 0 B) +7 C) +3 D) +4
- The oxidation numbers of all the atoms in  $\text{H}_2\text{SO}_4$  must add up to  
A) 0 B) +5 C) +9 D) +16
- In which compound is the oxidation number of oxygen treated as -1?  
A)  $\text{CO}$  B)  $\text{CO}_2$  C)  $\text{H}_2\text{O}$  D)  $\text{H}_2\text{O}_2$
- What is the oxidation number of iodine in  $\text{KIO}_4$ ?  
A) +1 B) -1 C) +7 D) -7
- What is the oxidation number of chromium in the chromate ion,  $\text{CrO}_4^{2-}$ ?  
A) +6 B) +2 C) +3 D) +8