1. Which is the oxidizing agent in the reaction	7. According to Reference Table J, which is the strongest reducing agent?		
$2 \text{ Fe}^{2+} + \text{Cl}_2 \rightarrow 2 \text{ Fe}^{3+} + 2 \text{ Cl}^-?$	A) Li(s) B) Na(s)		
A) Fe^{2+} B) Cl_2 C) Fe^{3+} D) Cl^-	C) $F_2(g)$ D) $Br_2(\ell)$		
2. In the reaction	8. A reducing agent is a substance that		
$Zn + Fe^{2+} \rightarrow Zn^{2+} + Fe$,	A) gains protonsB) loses protonsC) gains electronsD) loses electrons		
the reducing agent is	9. In a redox reaction, a reducing agent will always		
A) Zn B) Fe^{2+} C) Zn^{2+} D) Fe	A) lose electrons B) lose protons		
3. In the reaction	C) gain electrons D) gain protons		
Cu^0 + 2 H ₂ SO ₄ → $CuSO_4$ + 2 H ₂ O + SO ₂ ,	10. When a substance is oxidized, it		
	A) loses protons		
	B) gains protonsC) acts as an oxidizing agent		
copper is	D) acts as a reducing agent		
A) reduced and is the reducing agent B) reduced and is the reducing agent	11. As the elements in Period 3 of the Periodic Table are		
C) oxidized and is the oxidizing agent	considered in order of increasing atomic number, the ability of each successive element to act as a reducing agent		
D) oxidized and is the reducing agent			
4. According to Reference Table J, which species is the	A) decreases B) increases		
strongest oxidizing agent?	C) remains the same		
A) $Li(s)$ B) Li^+ C) $F_2(g)$ D) F^-	12. In a redox reaction, the species reduced		
5. In the reaction	A) gains electrons and is the oxidizing agent		
$Pb + 2 Ag^+ \rightarrow Pb^{2+} + 2 Ag$,	 B) gains electrons and is the reducing agent C) loses electrons and is the oxidizing agent D) loses electrons and is the reducing agent 		
the oxidizing agent is	D) loses electrons and is the reducing agent		
A) Ag^+ B) Ag C) Pb D) Pb^{2+}	13. In the equation:		
6. In the reaction	$Cu(s) + 2 Ag^{+}(aq) \leftrightarrow Cu^{2+}(aq) + 2 Ag(s)$		
$2 \text{ Mg} + \text{O}_2 \rightarrow 2 \text{ MgO},$	the oxidizing agent is		
the magnesium is the	A) Cu^0 B) Ag^+ C) Cu^{2+} D) Ag^0		
A) oxidizing agent and is reduced	14. In the reaction: $2H_2S + 3O_2 \rightarrow 2SO_2 + 2H_2O_3$		
B) oxidizing agent and is oxidized			
C) reducing agent and is reduced			
D) reducing agent and is oxidized	the oxidizing agent is		
	A) oxygen B) water		
	C) sulfur dioxide D) hydrogen sulfide		

15. Given the redox reaction:

 $Fe^{2+}(aq) + Zn(s) \rightarrow Zn^{2+}(aq) + Fe(s)$

Which species acts as a reducing agent?

A)	Fe(s)	B)	Fe ²⁺ (aq)
C)	Zn(s)	D)	$Zn^{2+}(aq)$

16. In a redox reaction, the reducing agent will

- A) lose electrons and be reduced
- B) lose electrons and be oxidized
- C) gain electrons and be reduced
- D) gain electrons and be oxidized
- 17. Given the redox reaction:

 $Mg(s) + CuSO_4(aq) \rightarrow MgSO_4(aq) + Cu(s)$

Which species acts as the oxidizing agent?

A) Cu(s)	B) $Cu^{2+}(aq)$
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C) Mg(s)

D) $Mg^{2+}(aq)$

18. Given the electrochemical cell reaction:

 $Zn(s) + Ni^{2+}(aq) \rightarrow Zn^{2+}(aq) + Ni(s)$

Which species is the reducing agent?

A) Zn B) Ni²⁺ C) Zn²⁺ D) Ni

19. Given the redox reaction:

$$2 \; Cr(s) + 3 \; Sn^{2+}(aq) \to 2 \; Cr^{3+}(aq) + 3 \; Sn(s)$$

Which species serves as the reducing agent?

A) Cr B) Sn^{2+} C) Cr^{3+} D) Sn

20. The oxidation number of a reducing agent can change from