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- The gram formula mass of NH_4Cl is
A) 22.4 g/mole B) 28.0 g/mole
C) 53.5 g/mole D) 95.5 g/mole
 - A 1.0-mole sample of krypton gas has a mass of
A) 19 g B) 36 g C) 39 g D) 84 g
 - What is the gram formula mass of Li_2SO_4 ?
A) 54 g B) 55 g C) 110 g D) 206 g
 - What is the gram formula mass of $\text{Ca}(\text{OH})_2$?
A) 29 g B) 34 g C) 57 g D) 74 g
 - What is the gram formula mass of K_2CO_3 ?
A) 138 g B) 106 g C) 99 g D) 67 g
 - What is the gram formula mass of $\text{Mg}(\text{ClO}_3)_2$?
A) 107 g B) 142 g C) 174 g D) 191 g
 - The sum of the atomic masses of the atoms in one molecule of $\text{C}_3\text{H}_6\text{Br}_2$ is called the
A) formula mass
B) isotopic mass
C) percent abundance
D) percent composition
 - What is the gram-formula mass of $(\text{NH}_4)_3\text{PO}_4$?
A) 112 g/mol B) 121 g/mol
C) 149 g/mol D) 242 g/mol
 - The molar mass of $\text{Ba}(\text{OH})_2$ is
A) 154.3 g B) 155.3 g
C) 171.3 g D) 308.6 g
 - What is the gram formula mass of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$?
A) 160. g B) 178 g
C) 186 g D) 250. g
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