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- All chemical bonds result from
 - overlapping of unoccupied orbitals of two or more atoms.
 - loss of one or more electrons by atoms which are gained by other atoms.
 - simultaneous attraction of one electron by two nuclei.
 - overlapping of two electron-filled orbitals with different energy levels.
 - the attraction of two ions.
 - Atoms within a molecule are primarily held together by which attraction force?
 - electrical
 - magnetic
 - gravitational
 - van der Waals
 - The sharing of electrons during bond formation always involves
 - formation of polar molecules.
 - formation of positive and negative ions.
 - shared electrons being attracted more by one atom than another.
 - lower energy content for the bonded atoms than for the unbonded atoms.
 - Which term indicates how strongly an atom attracts the electrons in a chemical bond?
 - alkalinity
 - atomic mass
 - electronegativity
 - activation energy
 - Electronegativity is a measure of an atom's ability to
 - attract the electrons in the bond between the atom and another atom
 - repel the electrons in the bond between the atom and another atom
 - attract the protons of another atom
 - repel the protons of another atom
 - Which term refers to how strongly an atom of an element attracts electrons in a chemical bond with an atom of a different element?
 - entropy
 - electronegativity
 - activation energy
 - first ionization energy
 - The amount of energy to remove its most loosely bound electron is the definition of:
 - electronegativity
 - ionization energy
 - chemical bond
 - radiation
 - Which element requires the *least* amount of energy to remove its most loosely bound electron?
 - Li
 - Mg
 - Ba
 - Ca
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