1. What is the most likely electronegativity value for a metallic element?	11. An element with an electronegativity of 0.9 bonds with an element with an electronegativity of 3.1. Which phrase best describes the bond between these elements?		
A) 1.3 B) 2.7 C) 3.4 D) 4.0			
2. Which element has the <i>lowest</i> electronegativity value?	 A) mostly ionic in character and formed between two nonmetals B) mostly ionic in character and formed between a metal and a nonmetal C) mostly covalent in character and formed 		
 A) F B) Fr C) Cl D) Cr 3. Which element has an atom with the greatest tendency to attract electrons in a chemical bond? 			
A) carbonB) chlorineC) siliconD) sulfur	between two nonmetalD) mostly covalent in character and formed between a metal and a nonmetal		
4. Which element has an atom with the greatest attraction for electrons in a chemical bond?	12. Electronegativity is a measure of an atom's ability to		
A) As B) Bi C) N D) P	A) attract the electrons in the bond between the atom and another atom		
5. Which term indicates how strongly an atom attracts the electrons in a chemical bond?	B) repel the electrons in the bond between the atom and another atom		
A) alkalinityB) atomic massC) electronegativityD) activation energy	C) attract the protons of another atomD) repel the protons of another atom		
6. Based on electronegativity values, which type of elements tends to have the greatest attraction for electrons in a bond?	13. Which type of bonding is usually exhibited when the electronegativity difference between two atoms is 1.1?		
A) metalsB) metalloidsC) nonmetalsD) noble gases	A) ionicB) covalentC) metallicD) network		
7. Which element has atoms with the greatest attraction	14. Which compound has the least ionic character?		
A) beryllium B) fluorine	A) KClB) CaCl2C) AlCl3D) CCl4		
C) lithium D) oxygen	15. Which compound contains a bond with the least		
8. Based on your Reference Tables, the atoms of which of these elements have the strongest attraction for	ionic character?		
electrons in a chemical bond?	16. Which pair of elements below will form a compound		
A) N B) Na C) P D) Pt	with the greatest ionic character?		
9. If the electronegativity difference between the elements in compound Na <i>X</i> is 2.1, what is element <i>X</i> ?	A) Pb and FB) Ca and OC) Na and ClD) Cs and N		
A) bromineB) chlorineC) fluorineD) oxygen	17. Which compound would have the greatest degree of ionic character?		
10. Which of the following elements has the greatest ability to attract electrons?	A) Na2O B) H2O C) CO2 D) NO2		
A) Li B) Be C) Na D) Mg			

18.	8. Which atom will form the most polar bond with the greatest degree of ionic bonding when bonding with sodium?			st polar bond with the ng when bonding with	19. As the difference in electronegativity between two atoms decreases, the tendency for the formation of covalent bonds	
	A) F	B) Cl	C) I	D) Br	A) decreasesC) remains the sa	B) increases me
			20. Which of the following compounds contains a bond with the greatest degree of ionic character?			
					A) CaO C) PH3	B) MgBr₂D) CCl₄