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1. The molecular formula of a given compound is C_4H_{10} . The empirical formula of the same compound would be
- A) C_2H_5 B) CH_2
C) C_3H_7 D) C_4H_{10}
2. The molecular formula of a given compound is $C_6H_{12}O_6$. The empirical formula of the same compound would be
- A) $C_{12}H_{24}O_{12}$ B) $C_3H_6O_3$
C) $C_6H_{12}O_6$ D) CH_2O
3. A substance has an empirical formula of CH_2 and a molar mass of 56 grams per mole. The molecular formula for this compound is
- A) CH_2 B) C_4H_6 C) C_4H_8 D) C_8H_4
4. The empirical formula of a compound is CH_3 . The molecular formula of this compound could be
- A) CH_4 B) C_2H_4 C) C_2H_6 D) C_3H_6
5. What is the molecular formula of a compound that has a molecular mass of 92 and an empirical formula of NO_2 ?
- A) NO_2 B) N_2O_4 C) N_3O_6 D) N_4O_8
6. What is the molecular formula of a compound that has a molecular mass of 42 and an empirical formula of CH_2 ?
- A) CH_2 B) C_2H_4 C) C_3H_6 D) C_4H_{12}
7. What is the molecular formula of a compound with an empirical formula of CH and a molecular mass of 78?
- A) C_6H_6 B) C_4H_{10}
C) C_2H_2 D) CH
8. A compound has a molar mass of 90. grams per mole and the empirical formula CH_2O . What is the molecular formula of this compound?
- A) CH_2O B) $C_2H_4O_2$
C) $C_3H_6O_3$ D) $C_4H_8O_4$
9. A compound has a gram formula mass of 56 grams per mole. What is the molecular formula for this compound?
- A) CH_2 B) C_2H_4 C) C_3H_6 D) C_4H_8
10. A compound has an empirical formula of HCO_2 and a molecular mass of 90. grams per mole. What is the molecular formula of this compound?
- A) HCO B) $H_2C_2O_4$
C) $H_4C_4O_8$ D) $H_6C_6O_{12}$
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