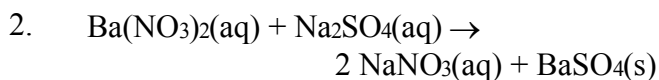


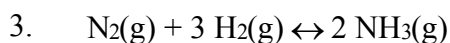
What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement



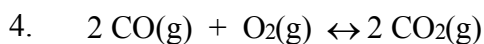
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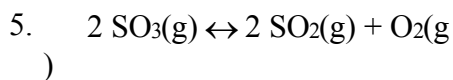
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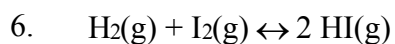
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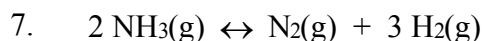
What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement



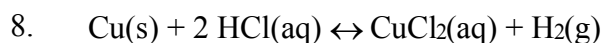
What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement



What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement

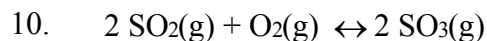


What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement

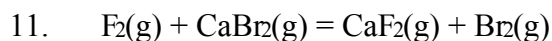
9. In which type of reaction do two or more substances combine to produce a single substance?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement



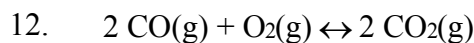
What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement



What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement



What type of reaction is shown above?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement

13. Which list includes three types of chemical reactions?

- A) decomposition, single replacement, and solidification
- B) decomposition, single replacement, and double replacement
- C) solidification, double replacement, and decomposition
- D) solidification, double replacement, and single replacement

14. Which equation represents a single replacement reaction?

- A) $2H_2O_2 \rightarrow 2H_2O + O_2$
- B) $2H_2 + O_2 \rightarrow 2H_2O$
- C) $H_2SO_4 + Mg \rightarrow H_2 + MgSO_4$
- D) $HCl + KOH \rightarrow KCl + H_2O$

15. Which change results in the formation of different substances?

- A) burning of propane
- B) melting of $NaCl(s)$
- C) deposition of $CO_2(g)$
- D) solidification of water

16. Which terms identify types of chemical reactions?

- A) decomposition and sublimation
- B) decomposition and synthesis
- C) deposition and sublimation
- D) deposition and synthesis

17. Given the word equation:

sodium chlorate \rightarrow sodium chloride + oxygen

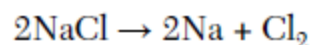
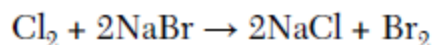
Which type of chemical reaction is represented by this equation?

- A) double replacement
- B) single replacement
- C) decomposition
- D) synthesis

18. In which type of chemical reaction do two or more reactants combine to form one product, only?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement

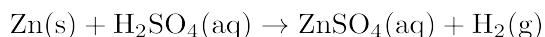
19. Given the balanced equations representing two chemical reactions:



Which type of chemical reactions are represented by these equations?

- A) single replacement and decomposition
- B) single replacement and double replacement
- C) synthesis and decomposition
- D) synthesis and double replacement

20. Given the balanced equation representing a reaction:



Which type of reaction is represented by this equation?

- A) decomposition
- B) double replacement
- C) single replacement
- D) synthesis

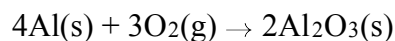
21. Which balanced equation represents a single-replacement reaction?

- A) $\text{Mg} + 2\text{AgNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + 2\text{Ag}$
- B) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
- C) $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
- D) $\text{MgCl}_2 + 2\text{AgNO}_3 \rightarrow 2\text{AgCl} + \text{Mg}(\text{NO}_3)_2$

22. Which equation represents a decomposition reaction?

- A) $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$
- B) $\text{Cu}(\text{s}) + 2\text{AgNO}_3(\text{aq}) \rightarrow 2\text{Ag}(\text{s}) + \text{Cu}(\text{NO}_3)_2(\text{aq})$
- C) $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
- D) $\text{KOH}(\text{aq}) + \text{HCl}(\text{aq}) \rightarrow \text{KCl}(\text{aq}) + \text{H}_2\text{O}(\text{l})$

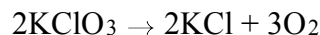
23. Given the balanced equation representing a reaction:



Which type of chemical reaction is represented by this equation?

- A) double replacement
- B) single replacement
- C) substitution
- D) synthesis

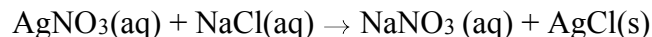
24. Given the balanced equation:



Which type of reaction is represented by this equation?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement

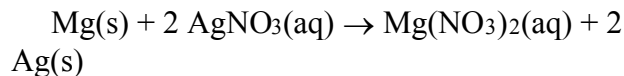
25. Given the balanced equation:



This reaction is classified as

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement

26. Given the reaction:



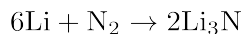
Which type of reaction is represented?

- A) single replacement
- B) double replacement
- C) synthesis
- D) decomposition

27. Which term identifies a type of chemical reaction?

- A) decomposition
- B) distillation
- C) sublimation
- D) vaporization

28. Given the balanced equation representing a reaction:



Which type of chemical reaction is represented by this equation?

- A) synthesis
- B) decomposition
- C) single replacement
- D) double replacement